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Note: # Neutrons = Mass Number - Atomic Number. # Neutrons of Nitrogen-14 =  $14 - 7 = 7$ . # Neutrons of Nitrogen-13 =  $13 - 7 = 6$ . c) The number of electrons in a neutral atom always equals the atomic number. The atomic number of

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both isotopes is 7, therefore the number of electrons in both isotopes is also 7.

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